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In aqueous solutions, protons do not exist in the free state, but rather as  $\text{H}_3\text{O}^+$  or  $\text{H}_5\text{O}_2^+$  ions but for simplicity one can refer to them as hydrogen ions or protons 2. Equilibrium constant of water dissociation reaction is  $K_{\text{eq}} = \frac{[\text{H}^+][\text{OH}^-]}{[\text{H}_2\text{O}]}$  3.

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1. Later reabsorption of energy by bond breaks it; C—C, C—H or C—O covalent bonds require 80 - 100 kcal/mole to break 2. This energy is quite large so these bonds are stable under most ...

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Section Reference: Section 2.5 Four Types of Biological Molecules Solution: Placement in a urea solution should totally denature the enzyme, especially since there are no disulfide linkages.

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